

XI Chemistry Worksheet

Time: 30 min

Ch # 7 : Equilibrium -02

Full Marks: 20

Instructions:

1. All questions are compulsory.
2. Please give the explanation for the answer where applicable.

Q1 - What are the applications of Equilibrium constant?

(3 Marks)

Q2 - What is Le Chatelier's principle?

(1 Mark)

Q3 - What are the effects of temperature, Pressure and concentration on the equilibrium?

(3 Marks)

Q4 - What are the effects of catalyst and inert gas addition?

(2 Marks)

Q5 - What is an Ionic Equilibrium?

(1 Mark)

Q6 -What is a chemical equilibrium?

(1 Mark)

Q7 -Define Law of chemical equilibrium.

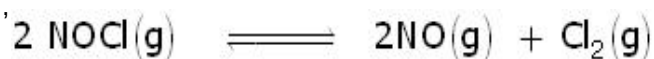
(2 Marks)

Q8 -(i)Define the term 'pH of solution'.

(ii) The hydrogen ion concentration of a solution is 10^{-4} . Calculate the pH of solution.

(2 Marks)

Q9 - For the equilibrium,



The value of equilibrium constant, K_c is 4.30×10^{-6} at 1069 K. calculate the K_p for the reaction at this temperature?

(3 Marks)

Q10 -What is Solubility Product?

(2 Marks)