

ASSIGNMENT CLASS – X CHEMISTRY LESSON - 1

22.05.14

Q1. Write two examples of everyday life, where redox reactions are taking place.

Q2. Food items can be preserved for longer time when kept in a refrigerator. What is the chemical reason behind this fact?

Q3. A student has been collecting silver coins and copper coins. One day she observed a black coating on silver coins and green coating on copper coins. Which chemical phenomenon is responsible for these coatings? Write the chemical name of black and green coatings.

Q4. Compound “A” when dissolved in water gives compound “B” and liberates heat. Compound “A” is used in white washing. Compound “B” reacts with CO₂ to form a white precipitate of compound “C” Identify compound A,B,C.

Q5. Shyam visited a government dispensary. He saw the medicine in dark bottles were not stored properly. They were not kept away from light and heat. Shyam reported the matter to the medical superintendent and ensured that all medicines restored properly. (a) Why are some medicines stored in cool places in dark bottles? (b) Why do some medicines need refrigeration? (c) What value of shyam is displayed?

Q6. Alka and Sudha were in the marriage procession of their brother. Fireworks were being used freely. The rocket were producing dazzling white light in the sky. Alka asked Sudha how the sparkles and dazzling

light are produced? Sudha explained the phenomenon and Alka was satisfied. (a) Which component of fireworks produces white dazzling light? (b) Is it a physical or chemical reaction? Give the reaction. (c) What values are shown by Sudha?

Q7. Explain all types of reactions with examples.

Q8. Translate the chemical reactions into equation then balance these

(a) Sodium hydroxide solution is treated with acetic acid to form sodium acetate and water.

(b) A
solution of potassium chloride when mixed with silver nitrate solution, an insoluble white substance of silver chloride and solution of potassium nitrate is formed.

(c) Copper sulphate on treatment with potassium iodide gives precipitates of cuprous iodide (Cu_2I_2) and liberates iodine gas as well as forms potassium sulphate.

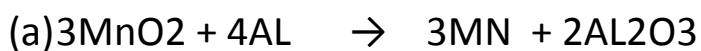
(d) An
aqueous solution of barium chloride is reacted with solution of sodium sulphate form white ppt. of barium sulphate and sodium chloride in solution.

Q9. Define rancidity and corrosion with example.

Q10. How can a balance equation be made more informative? Explain with example.

Q11. A white salt on heating decomposes to give brown fumes and residue is left behind. (i) Name the salt. (ii) Write the equation.

Q12. Name the substance oxidized and reduced in the given reactions :



Q13. Give reason why? (a) Decomposition reaction is called the opposite of combination reaction. (b) the respiration is considered an exothermic reaction. (c) magnesium ribbon should be cleaned before burning in air. (d) The colour of copper sulphate solution change when iron nail is dipped in it.

Q14. Explain exothermic and endothermic reactions with examples.

Q15. Grapes hanging on the plant do not ferment but after being plucked from the plant can be fermented. Under what conditions do these grapes ferment ? is it a chemical or a physical change ?