

ASSIGNMENT

Q1 Explain the following terms

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| (a) Neutralisation reaction | (b) Precipitation reaction |
| (c) Corrosion | (d) Rancidity |
| (e) Indicators (olfactory indicator) | (f) Ductility |
| (g) Malleability | (h) Sonorous |
| (i) Lustre | (j) Photochemical Reaction |
| (k) Exothermic & endothermic | (l) Water of crystallisation |
| (m) Oxidation & Reduction reaction | |

Q2 Define

- (a) Combination reaction
- (b) Decomposition reaction
- (c) displacement reaction
- (d) Double displacement reaction
- (e) Strong acids & bases
- (f) weak acids & bases
- (g) Aqua regia
- (h) Reactivity series (learn the series also)

Q3 Give reason Why:

- (a) Respiration is considered as exothermic reaction
- (b) Curd & sour substance not kept in brass vessel.
- (c) Dry HCl gas does not change the colour of dry litmus.
- (d) P.O.P. should be stored in moisture proof container.
- (e) Sodium kept in kerosene.
- (f) Ionic compounds have high melting point.

Q4 What are amphoteric oxides ? Give two examples with equations.

Q5 Differentiate between

- 1 Calcination & roasting
- 2 Mineral & ores
- 3 Metals & nonmetals
- 4 Properties of acids & bases

Q6 Explain electrolytic refining with dig.

Q7 Explain Reduction on the basis of position of metals on reactivity series.

Q8 Explain the activity to show that hydrogen gas is evolved when metal reacts with acids with dig.

Q9 Write the chemical name ,formula ,preparation, properties, & uses of
(a) Washing soda (b) Baking soda (c) Bleaching powder (d) plaster of paris.

Q10 Preparation of NaOH is known as chlor alkali process. explain.