

IMPORTANT COMPOUNDS AND THEIR FORMULAE

1. Active nitrogen $\rightarrow$ Atomic nitrogen
2. Alums $\rightarrow$ $MAl(SO_4)_2 \cdot 12H_2O$; (M = NH_4^+ , Na^+ , K^+ etc.)
3. Amatol $\rightarrow$ 80% NH_4NO_3 + 20% T.N.T. (explosive)
4. Anhydron $\rightarrow$ $Mg(ClO_4)_2$
5. Aqua regia $\rightarrow$ conc. $1HNO_3$ + conc. $3HCl$
6. Arsine $\rightarrow$ AsH_3
7. Asbestos $\rightarrow$ $CaMg_3(SiO_3)_4$
8. Borane $\rightarrow$ Hydrides of Boron
9. Bremstone $\rightarrow$ S_8
10. Blue vitriol $\rightarrow$ $CuSO_4 \cdot 5H_2O$
11. Bleaching powder $\rightarrow$ $Ca(OCl)Cl$
12. Baryta water $\rightarrow$ $Ba(OH)_2$ Solution
13. Baryta $\rightarrow$ BaO

14. Baking powder or Soda → NaHCO_3
15. Black jack → Zinc ore
16. Calgon → $\text{Na}_2[\text{Na}_4(\text{PO}_3)_6]$
17. Carborundum → SiC
18. Caliche → Natural NaNO_3
containing NaIO_3
19. Caustic Soda → NaOH
20. Caustic potash → KOH
21. Calomel → Hg_2Cl_2
22. Cerussite → PbCO_3
23. Cementite → Fe_3C (iron carbide)
24. Chrom alum → $\text{K}_2\text{SO}_4 \cdot \text{Cr}_2(\text{SO}_4)_3 \cdot 24\text{H}_2\text{O}$
25. Chinese white → ZnO
26. Corrosive sublimate → HgCl_2
27. D.D.T. → p-dichloro-diphenyl-trichloro-ethene
28. Deuterium → D or ${}_1\text{H}^2$ (Isotope of hydrogen)
29. Dry ice → Solid CO_2
30. Fehling's solution → A deep blue solution = $\text{CuSO}_4 \cdot 5\text{H}_2\text{O} + \text{NaOH} + \text{Na, K-tartrate}$ (used for the test of aldehydes)
31. Feldspar → KAlSi_3O_8

32. Fenton's reagent $\rightarrow$ H_2O_2 + few drops of FeCl_3
33. Freon $\rightarrow$ CF_2Cl_2
34. Ferric alum $\rightarrow$ $\text{K}_2\text{SO}_4 \cdot \text{Fe}_2(\text{SO}_4)_3 \cdot 24\text{H}_2\text{O}$
35. Fusion mixture $\rightarrow$ $\text{Na}_2\text{CO}_3 + \text{K}_2\text{CO}_3$
36. Fluid magnesia $\rightarrow$ 12% aqueous soln. of $\text{Mg}(\text{HCO}_3)_2$
37. Glauber's salt $\rightarrow$ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$
38. Graphite $\rightarrow$ An allotrope of carbon
39. Green vitriol $\rightarrow$ $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$
40. Gun powder $\rightarrow$ 75% KNO_3 , 12% S, 13% charcoal
41. Heavy hydrogen $\rightarrow$ D_2
42. Heavy water $\rightarrow$ D_2O
43. Hydrolith $\rightarrow$ CaH_2
44. Hypo $\rightarrow$ $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$
45. Killed spirit $\rightarrow$ $\text{ZnCl}_2 + \text{ZnO}$ (Zn-oxy chloride)
46. Kesserite $\rightarrow$ $\text{MgSO}_4 \cdot \text{H}_2\text{O}$
47. Leuna saltpetre $\rightarrow$ Fertilizer $[\text{NH}_4\text{NO}_3 + (\text{NH}_4)_2\text{SO}_4]$
48. Lime or Quick lime $\rightarrow$ CaO
49. Lead of pencil $\rightarrow$ Graphite (C)

50. Lime water → A clear aqueous solution of Ca(OH)_2
51. Laughing gas → N_2O
52. Lunar caustic → AgNO_3
53. Litharge → PbO
54. Lithopone → A white pigment ($\text{ZnS} + \text{BaSO}_4$)
55. Massicot → PbO
56. Matte → $\text{Cu}_2\text{S} + \text{FeS}$
57. Magnesia alba → $2\text{MgCO}_3 \cdot \text{Mg(OH)}_2 \cdot 3\text{H}_2\text{O}$
58. Magnesia → MgO
59. Marsh gas → Methane (CH_4)
60. Marble → CaCO_3
61. Micro cosmic salt → $\text{NaNH}_4\text{HPO}_4$ (used in the test of silicates)
62. Milk of magnesia → A paste of Mg(OH)_2 in water
63. Mohr's salt → $\text{FeSO}_4(\text{NH}_4)_2\text{SO}_4 \cdot 6\text{H}_2\text{O}$
64. Muriatic acid → HCl
65. Milk of lime → Suspension of Ca(OH)_2 in water
66. Minium → Pb_3O_4
67. Nascent hydrogen → Atomic hydrogen
68. Nessler's reagent → Aq. soln. of K_2HgI_4

69. Nitro chalk	→ Fertilizer $[\text{NH}_4\text{NO}_3 + (\text{NH}_4)_2\text{CO}_3]$
70. Nitrolim	→ CaCN_2
71. Nitrophos	→ $\text{Ca}(\text{H}_2\text{PO}_4)_2 + 2\text{Ca}(\text{NO}_3)_2$
72. Oil of vitriol	→ conc. H_2SO_4
73. Ozone	→ O_3
74. Oleum	→ $\text{H}_2\text{S}_2\text{O}_7$
75. Oxygen gas	→ O_2
76. Pharaoh's Serpents	→ $\text{Hg}(\text{CNS})_2$
77. Philosopher's Wool	→ ZnO
78. Phosphine	→ PH_3
79. Phosgene	→ COCl_2
80. Pig iron	→ impure form of iron
81. Potas alum	→ $\text{K}_2\text{SO}_4 \cdot \text{Al}_2(\text{SO}_4)_3 \cdot 24\text{H}_2\text{O}$
82. Producer gas	→ A mixture of $(\text{CO} + \text{N}_2)$
83. Plaster of paris	→ $\text{Ca}(\text{SO}_4)_2 \cdot \text{H}_2\text{O}$
84. Quartz	→ $(\text{SiO}_2)_n$
85. Quick silver	→ Hg
86. Quick lime	→ CaO
87. Red lead	→ Pb_3O_4
88. Refrigerant	→ $\text{NH}_3, \text{CO}_2, \text{CF}_2\text{Cl}_2$ etc.
89. Rochelle Salt	→ Sodium potassium tartrate

90. Rust $\rightarrow \text{Fe}_2\text{O}_3 \cdot x\text{H}_2\text{O}$
91. Sorel's Cement $\rightarrow \text{Mg}(\text{OH})\text{Cl}$
92. Soda-lime $\rightarrow \text{NaOH} + \text{CaO}$
93. Soda ash or Sal soda $\rightarrow \text{Na}_2\text{CO}_3$
94. Spathose ore $\rightarrow \text{FeCO}_3$
95. Salammoniac $\rightarrow \text{NH}_4\text{Cl}$
96. Slaked lime $\rightarrow \text{Ca}(\text{OH})_2$
97. Sal volatile (smelling salt) $\rightarrow (\text{NH}_4)_2\text{CO}_3$
98. Spinel $\rightarrow \text{MgAl}_2\text{O}_4$
99. Superphosphate $\rightarrow \text{Ca}(\text{H}_2\text{PO}_4)_2 + 2\text{CaSO}_4$
100. T.N.T. $\rightarrow$ Tri-nitro toluene (explosive)
101. T.N.B. $\rightarrow$ Tri-nitro benzene (more powerful explosive than T.N.T.)
102. Tincal $\rightarrow \text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$
103. Talc $\rightarrow 3\text{MgO} \cdot 4\text{SiO}_2 \cdot \text{H}_2\text{O}$
104. Tritium $\rightarrow$ T or, H^3 , an isotope of hydrogen
105. Vermilion $\rightarrow \text{HgS}$ (red)
106. Water glass $\rightarrow$ Sodium metasilicate (Na_2SiO_3)
107. Water gas $\rightarrow \text{CO} + \text{H}_2$

108. Wrought iron → Pure form of iron
109. White vitriol → $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$
110. White lead → $2\text{PbCO}_3 \cdot \text{Pb(OH)}_2$
111. Zinc white → ZnO
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